

Sci Method and Measuring Warm Up

1. This is the variable that is changed during an experiment
2. This is the variable that is measured during an experiment
3. Data points that are very close to each other, but not close to the accepted value (standard) could be described as being:
4. mL stands for _____ and measures _____

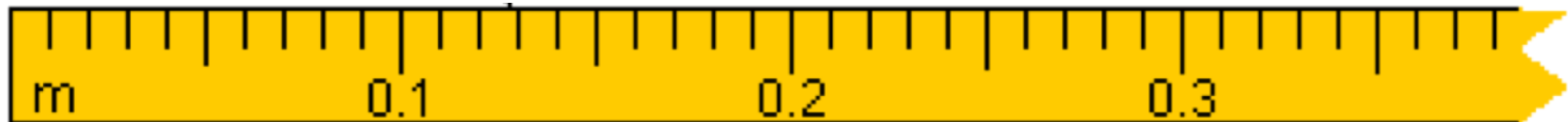
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5. Precisely record the following measurement of letter "A"

6. Precisely record the following measurement of letter "B"

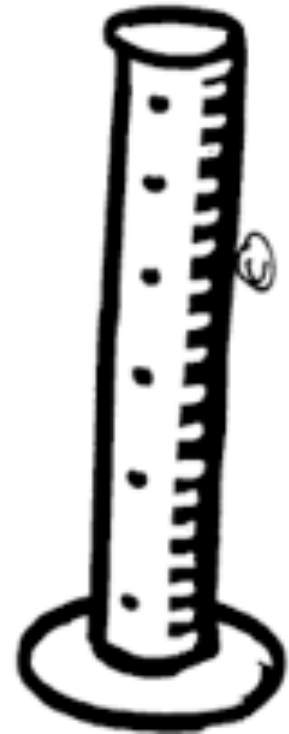
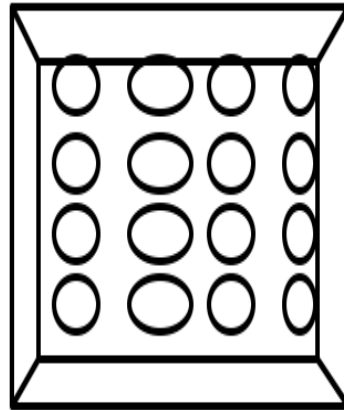
7. Precisely record the following measurement of letter "C"

A B C
↓ ↓ ↓



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8. Name these pieces of equipment



Matter Warm Up

1. Classify the following as : element, compound, mixture, pure substance, heterogeneous mixture, homogeneous mixture
 1. Salt : NaCl
 2. Salt Water: NaCl + H₂O
 3. Sodium: Na
2. True or false: oxidation is a physical property
3. How many of each type of atom are present in the following formula: H₃PO₄
4. Melting is a physical _____.
5. Corrosion is a _____ property.

Atomic Structure Warm Up

1. Identify the number of protons, neutrons, and electrons in a neutral atom of C-14. Rewrite using another form of isotope notation
2. The scientist accredited with developing the plum pudding model is:
3. The _____ experiment demonstrated that the nucleus has a positive charge
4. The atomic number of an atom with 10 protons and 10 neutrons is:
5. The mass number of an atom with 10 protons and 10 neutrons is:

Nuclear Warm Up

1. During nuclear decay the _____ of the atom releases _____
2. During this type of decay, the atomic number increases by 1
3. During this type of decay, the atomic number decreases by 2
4. The purpose of the control rods in a nuclear reactor is.....
5. Calculate the percent remaining of a sample that has a half life of 5 hrs, after 25 hrs
6. Compare and contrast fission and fusion

Electrons Warm Up

1. Electron A falls from $n=5 \rightarrow n=2$. Electron B falls from $n=3 \rightarrow n=2$. One produces a blue color light while the other produces a red color light. Which is which?
2. How many electrons can fit in a 4p sublevel?
3. How many electrons can fit in a 3d orbital?
4. Compare and contrast a 3p and 4s orbital
5. Write the electron configuration for Hg
6. An element in group 2 has valence electrons in what sublevel?