

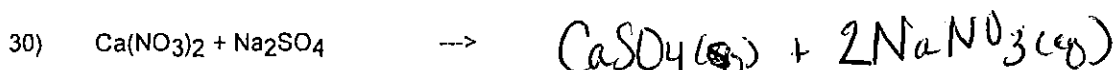
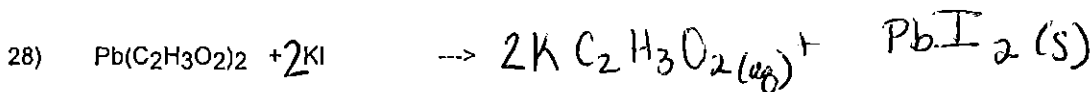
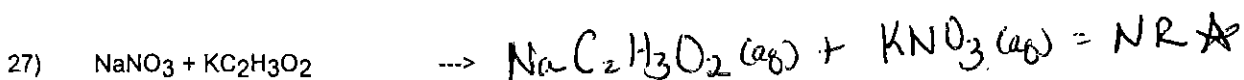
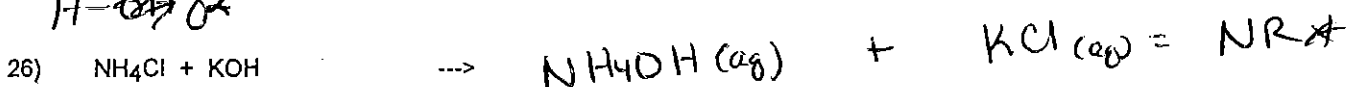
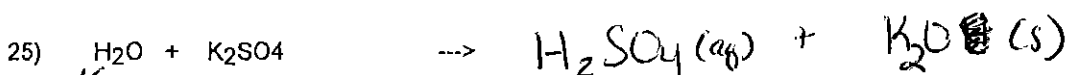
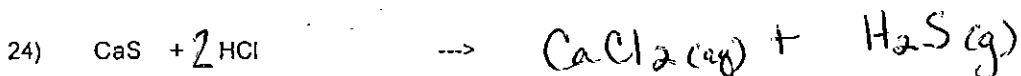
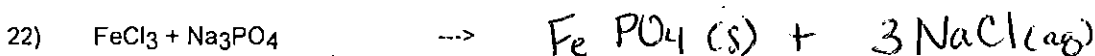
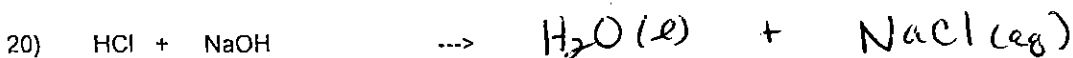
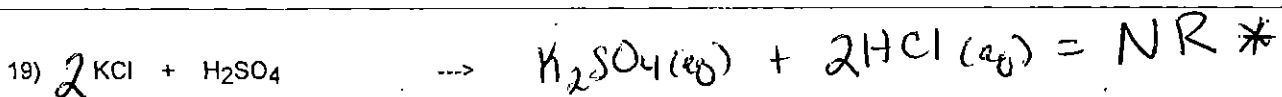
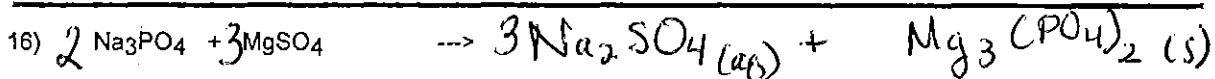
CHEMISTRY **DOUBLE REPLACEMENT REACTION WORKSHEET**

PRACTICE REACTIONS

- 1) $3\text{Ca(OH)}_2 + 2\text{H}_3\text{PO}_4 \rightarrow \text{Ca}_3(\text{PO}_4)_2(\text{s}) + 6\text{H}_2\text{O}(\text{l})$
- 2) $\text{K}_2\text{CO}_3 + \text{BaCl}_2 \rightarrow 2\text{KCl}(\text{aq}) + \text{BaCO}_3(\text{s})$
- 3) $\text{Cd}_3(\text{PO}_4)_2 + 3(\text{NH}_4)_2\text{S} \rightarrow 3\text{CdS}(\text{s}) + 2(\text{NH}_4)_3\text{PO}_4(\text{aq})$
- 4) $\text{Co(OH)}_3 + 3\text{HNO}_3 \rightarrow \text{Co}(\text{NO}_3)_3(\text{aq}) + 3\text{H}_2\text{O}(\text{l})$
- 5) $\text{AgNO}_3 + \text{KCl} \rightarrow \text{AgCl}(\text{s}) + \text{KNO}_3(\text{aq})$
- * 6) $\text{Na}_2\text{CO}_3 + \text{H}_2\text{SO}_4 \rightarrow \text{Na}_2\text{SO}_4(\text{aq}) + \text{H}_2\text{CO}_3 \rightarrow \text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{l})$
- 7) $\text{Al(OH)}_3 + 3\text{HC}_2\text{H}_3\text{O}_2 \rightarrow \text{Al}(\text{C}_2\text{H}_3\text{O}_2)_3(\text{aq}) + 3\text{H}_2\text{O}(\text{l})$
- 8) $\text{Al}_2(\text{SO}_4)_3 + \text{Ca}_3(\text{PO}_4)_2 \rightarrow \text{AlPO}_4(\text{s}) + \text{CaSO}_4(\text{aq})$
- * 9) $\text{Cr}_2(\text{SO}_3)_3 + 3\text{H}_2\text{SO}_4 \rightarrow \text{Cr}_2(\text{SO}_4)_3(\text{aq}) + 3\text{H}_2\text{SO}_3 \rightarrow \text{SO}_2(\text{g}) + \text{H}_2\text{O}(\text{l})$
- 10) $2\text{AgC}_2\text{H}_3\text{O}_2 + \text{K}_2\text{CrO}_4 \rightarrow \text{Ag}_2\text{CrO}_4 + 2\text{K}(\text{C}_2\text{H}_3\text{O}_2)(\text{aq})$
- 11) $\text{FeBr}_2 + \text{K}_2\text{CO}_3 \rightarrow \text{FeCO}_3(\text{s}) + 2\text{KBr}(\text{aq})$
- 12) $\text{Ag}_2\text{S} + \text{CuCl}_2 \rightarrow 2\text{AgCl}(\text{s}) + \text{CuS}(\text{s})$
- 13) $\text{Pb}(\text{NO}_3)_2 + 2\text{HI} \rightarrow 2\text{HNO}_3(\text{aq}) + \text{PbI}_2(\text{s})$
- 14) $\text{Ba}(\text{ClO}_3)_2 + \text{H}_2\text{SO}_4 \rightarrow 2\text{HClO}_3(\text{aq}) + \text{BaSO}_4(\text{s})$
- * 15) $\text{CuS} + 2\text{KCl} \rightarrow \text{CuCl}_2(\text{aq}) + \text{H}_2\text{S}(\text{aq}) \rightarrow \boxed{\text{NR}}$

CHEMISTRY

DOUBLE REPLACEMENT REACTION WORKSHEET



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