

MIXED

The two lists below are a mixture of ionic and covalent compounds. First, tell whether the formula/name represents an ionic or covalent compound. Then, write the appropriate formula/name on the second line.

Remember:

Ionic compounds = start with a metal element or ammonium (NH_4^+)

Covalent compounds = start with a nonmetal or semimetal

	Type of Compound	Name Formula
1. CS_2	covalent	carbon disulfide
2. PbCO_3	ionic	lead (II) carbonate
3. $\text{K}_2\text{Cr}_2\text{O}_7$	ionic	potassium dichromate
4. $\text{Cd}(\text{NO}_2)_2$	ionic	cadmium (II) nitrite
*5. As_2O_3	Covalent - hydrous	diarsenic trioxide
6. $\text{Fe}(\text{OH})_3$	Iron ionic	Iron (III) hydroxide
7. Ag_3PO_4	ionic	Silver (I) phosphate
8. KCN	ionic	potassium cyanide
9. XeF_4	covalent	Xenon tetrafluoride
10. NaHCO_3	ionic	Sodium bicarbonate

	Type of Compound	FORMULA Name
1. strontium carbonate	ionic	SrCO_3
2. lithium sulfide	ionic	Li_2S
3. nitrogen trichloride	covalent	NCl_3
4. copper (I) sulfate	ionic	Cu_2SO_4
5. triphosphorous tetrasulfide	Covalent	P_3S_4
6. iodine pentafluoride	covalent	IF_5
7. lead (II) acetate	ionic	$\text{Pb}(\text{C}_2\text{H}_3\text{O}_2)_2$
8. phosphorous pentabromide	covalent	PBr_5
9. calcium nitride	ionic	Ca_3N_2
10. strontium hydroxide	ionic	$\text{Sr}(\text{OH})_2$